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[K]= [° C]+273.15 Another absolute

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temperature scale – the Rankine scale – can be used to convert between SI and English systems:

$[^{\circ}R] = 1.8 [K]$ $[^{\circ}R] = [^{\circ}F] + 459.67$ Of course, one can then write that $[^{\circ}F] = 1.8 [^{\circ}C] + 32$

1.2.3 PRESSURE

Pressure, p , is also an intensive property.

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Many of the definitions below are also used in the thermodynamics of chemical reactions. General basic quantities. Quantity (Common

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Name/s	(Common)	Symbol/s	SI Units
Dimension	Number of molecules	N	dimensionless
Number of moles	n	mol	...
...	C	V	$= \quad / \quad J K^{-1}$
[M]	[L	...	

Table of thermodynamic equations -
Wikipedia

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In thermodynamics and chemical engineering, the vapor–liquid equilibrium (VLE) describes the distribution of a chemical species between the vapor phase and a liquid phase.. The concentration of a vapor in contact with its liquid, especially at equilibrium, is often expressed in

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terms of vapor pressure, which will be a partial pressure (a part of the total gas pressure) if any other gas(es ...

Vapor-liquid equilibrium - Wikipedia
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Thermodynamics K.V. NARAYANAN
Former Professor and Head
Department of Chemical Engineering
and Former Principal Government
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on the most important
thermodynamic formulas and
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